

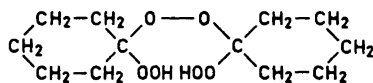
Crystal Structure of 1,1'-Dihydroperoxy- cyclohexanylperoxide-1,1'

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The crystals are monoclinic with cell dimensions $a=10.50_4$ Å, $b=6.99_1$ Å, $c=19.08_8$ Å and $\beta=107.7_8^\circ$. The space group is $P2_1/c$ and the cell contains four molecules. The phase problem was solved three-dimensionally by means of a computer procedure based on direct methods. Full-matrix least squares refinement gave the R -value of 8.6 % for 1637 observed reflections. Rigid-body motion analysis including the S-tensor has been carried out. Mean values of O—O, C—O, and C—C distances are 1.47₇ Å, 1.43₇ Å, and 1.53₃ Å, respectively. The average C—O—O angle for the hydroperoxy groups is 110.4° and the corresponding value for the peroxy bridge 107.9°. The asymmetric environments of the non-methylene carbon atoms resemble that observed for trimeric acetone peroxide, and may possibly be related to *intra*-molecular oxygen-hydrogen repulsions. Weak *inter*-molecular (2.81₇ Å) and *intra*-molecular (2.90₄ Å) hydrogen bonds are present, the *inter*-molecular bonds forming chains along b -axis. The conformation of the molecule is *cis*, and the dihedral angle C—O—O—C is 126.3°.

1,1'-Dihydroperoxycyclohexanylperoxide-1,1'



has been synthesized by T. Ledaal at this university by oxidation of cyclohexanone peroxide by means of hydrogen peroxide in the presence of HClO_4 . No crystal structure determinations of compounds containing hydroperoxy groups have been reported. Structural information concerning covalent organic peroxides is also limited.¹⁻⁴ The crystal structure analysis of this compound was carried out in order to obtain information about the hydroperoxy group, specially its ability to form hydrogen bonds, and to determine the dihedral angle of the peroxy group.

CRYSTAL DATA

The crystals belong to the monoclinic system and the systematic absences lead to the space group $P2_1/c$. The cell parameters, measured by means of a four circle diffractometer, and their estimated standard deviations are:

$$\begin{array}{ll} a = 10.504 \text{ \AA} & \sigma(\text{\AA}) = 0.003 \\ b = 6.991 \text{ \AA} & \sigma(\text{\AA}) = 0.002 \\ c = 19.088 \text{ \AA} & \sigma(\text{\AA}) = 0.002 \\ \beta = 107.78^\circ & \sigma(^\circ) = 0.01 \end{array}$$

The unit cell contains four molecules ($\rho_{\text{calc}}=1.30 \text{ gcm}^{-3}$, $\rho_{\text{obs}}=1.28 \text{ gcm}^{-3}$).

Intensity data were obtained by photometric measurements of integrated Weissenberg diagrams corresponding to $0kl$, $h0l$, . . . , $h4l$ ($\text{CuK}\alpha$ -radiation). The diagrams were taken at room temperature, and due to instability four crystals were used. 1637 independent reflections were strong enough to be measured while 309 were given $(1/4)I_{\text{min}}$ -values.

After cross-layer scaling the intensities were statistically put on absolute scale, and the overall temperature factor, $B=3.85 \text{ \AA}^2$, was used in the calculation of unitary structure factors, and as starting parameter in the isotropic refinement.

No corrections have been made for absorption or secondary extinction effects.

DETERMINATION OF THE STRUCTURE

Solution of the phase problem was obtained by applying a set of FORTRAN IV programs for CDC 3300⁵ based on direct methods.⁶⁻⁸ As to be expected from the $|U|$ -distribution given in Table 1, the programmed systematic

Table 1. Distribution of $|U|$ -values.

Range of $ U $	Number of reflections
0.00—0.01	62
0.01—0.02	281
0.02—0.03	97
0.03—0.04	138
0.04—0.05	102
0.05—0.06	110
0.06—0.07	117
0.07—0.08	103
0.08—0.09	111
0.09—0.10	88
0.10—0.15	378
0.15—0.20	190
0.20—0.25	96
0.25—0.30	31
0.30—0.35	21
0.35—0.40	7
0.40—0.50	4
0.50—1.00	0

application of Harker-Kasper inequalities gave no information about signs. A program, based on Sayre's sign relationship, which allows introduction of eight symbols for unknown signs, was used in deriving about 250 signs in terms of three unknowns. The corresponding eight three-dimensional Fourier syntheses were calculated and scanned for peaks larger than $2 \text{ e} \cdot \text{\AA}^{-3}$ followed by calculations of all distances less than 2 \AA between peak maxima. The resulting eight lists of coordinates and distances were examined, and one of them contained the expected number of peaks, with roughly the same heights, and with reasonable distances between peak maxima.

REFINEMENT

The R -factor corresponding to the coordinates thus obtained and the statistically determined overall temperature factor was $R=43.2\%$ for the 500 largest structure factors. Using a program based on the minimum residual method,⁹ this R -value was reduced to 14.2% .

A full-matrix least squares refinement⁵ was carried out. The weighting scheme No. 1:

$$\begin{array}{ll} \text{for } F_o \leq \text{FB}, & W = A1(F_o)^{B1} \\ \text{for } F_o > \text{FB}, & W = A2(F_o)^{B2} \end{array}$$

was adapted by taking:

$$A1 = 10.0, \quad A2 = 12.2, \quad B1 = 0.0, \quad B2 = -0.5 \quad \text{and} \quad \text{FB} = 1.5.$$

The atomic form factors used were those of Hanson *et al.*¹⁰ After two cycles of isotropic refinement, anisotropic thermal vibrational parameters were introduced for the oxygen and carbon atoms and three least squares cycles calculated. Positions of the 20 methylene hydrogen atoms were calculated assuming tetrahedral C—H bonds of length 1.03 \AA . The scale factors of the individual layers were adjusted by structure factor calculations. An attempt to localize the hydroperoxy hydrogen atoms from a difference fourier map was unsuccessful. Isotropic temperature factors, $B = 5.0 \text{ \AA}^2$, were applied for the methylene hydrogens, and their positional parameters only were refined. The final R -value for 1637 observed reflections was 8.6% .

309 accidentally absent reflections had previously been given $(1/4)I_{\text{min}}$ -values and had been scaled along with the rest of the data. Allowance for inclusion of these reflections in the refinement was made by applying weighting scheme No. 3 ($W=1/\sigma$). Satisfactory weight analysis was obtained by giving all unobserved reflections the same constant weight $W=A1/3$. The R -factor for all data was 9.5% .

When checking the 250 signs originally determined no discrepancies were found.

Table 2. Fractional atomic coordinates for oxygen and carbon atoms with estimated standard deviations (multiplied by 10^6).^a

Atom	<i>x</i>	<i>y</i>	<i>z</i>
O ₁	25570 28	19558 53	25861 17
O ₂	30066 26	38712 47	28599 15
O ₃	39499 23	27596 47	40646 15
O ₄	53078 22	34336 46	41080 14
O ₅	54315 23	14989 45	31287 13
O ₆	55725 27	31622 53	26875 16
C ₁	30194 34	40749 69	36147 20
C ₂	16856 38	34926 82	37194 28
C ₃	05744 37	48788 85	33164 27
C ₄	09244 42	69300 86	35557 27
C ₅	22294 44	75208 87	34130 29
C ₆	33704 39	61562 74	37885 25
C ₇	60092 31	18974 67	38956 20
C ₈	74379 33	26463 78	40839 24
C ₉	83658 38	10553 86	39541 27
C ₁₀	82804 38	-07473 86	43882 26
C ₁₁	68472 37	-15237 81	41671 25
C ₁₂	58901 35	00160 71	42755 22

^a For numbering of atoms, see Fig. 1.

Table 3. Anisotropic thermal vibration parameters and their estimated standard deviations (multiplied by 10^5).

Atom	B_{11}	B_{22}	B_{33}	B_{12}	B_{13}	B_{23}
O ₁	969 27	2045 97	428 10	-368 84	257 26	-366 49
O ₂	876 24	1991 93	338 8	-16 75	412 22	6 42
O ₃	666 21	2064 97	362 8	194 70	415 21	286 42
O ₄	636 21	1819 92	371 8	114 68	329 21	-175 41
O ₅	775 22	2017 91	292 7	107 70	251 21	-11 39
O ₆	1037 27	2880 108	350 9	620 88	541 25	606 49
C ₁	756 31	1789 134	307 10	494 103	333 29	121 58
C ₂	776 34	2113 152	470 15	451 110	564 36	433 72
C ₃	747 33	2568 168	478 15	229 121	479 36	338 80
C ₄	940 36	2525 170	418 14	977 128	434 37	205 75
C ₅	1006 40	1975 139	489 16	570 121	473 41	81 76
C ₆	869 34	1630 133	419 13	79 109	476 34	-243 65
C ₇	646 28	1571 131	315 10	198 92	238 27	-84 55
C ₈	638 30	2172 155	397 13	-87 102	308 31	-198 67
C ₉	709 32	2857 177	454 14	375 112	423 34	321 78
C ₁₀	793 33	2475 168	461 15	956 117	430 35	441 76
C ₁₁	855 34	2026 137	397 13	547 109	367 34	61 65
C ₁₂	747 29	1770 130	369 11	86 104	362 30	49 61

The fractional atomic coordinates and the thermal vibration parameters arrived at are given in Tables 2, 3, and 4; the expression for the anisotropic vibrations being:

$$\exp[-(B_{11}h^2 + B_{22}k^2 + B_{33}l^2 + B_{12}hk + B_{13}hl + B_{23}kl)]$$

Table 4. Fractional coordinates (multiplied by 10^3) for hydrogen atoms. The estimated standard deviations range from 0.003 to 0.01.^a

Atom	<i>x</i>	<i>y</i>	<i>z</i>
H2e	149	215	355
H2a	185	356	424
H3e	-039	449	338
H3a	028	477	271
H4e	028	790	333
H4a	110	704	410
H5e	246	882	364
H5a	214	748	284
H6e	417	625	365
H6a	369	629	434
H8e	740	388	378
H8a	772	306	463
H9e	936	149	418
H9a	810	078	337
H10e	891	-196	434
H10a	872	-027	499
H11e	688	-278	447
H11a	657	-213	363
H12e	504	-047	408
H12a	614	037	481

^a Referring to Fig. 1 the hydrogen atoms Hne and Hna are, respectively, equatorially and axially bonded to Cn.

In Table 5 calculated structure factors are compared with the observed values (and unobserved with $F_u = \frac{1}{2} F_o$ (min)).

THERMAL MOTIONS

The principal axes of the thermal vibration ellipsoids for oxygen and carbon atoms were calculated from the temperature parameters given in Table 3. Root mean square amplitudes and the corresponding *B*-values for the atomic anisotropic thermal vibration along the principal axes together with the components of these axes along the crystal axes are given in Table 6. The thermal vibration parameters of Table 3 were applied in a rigid-body analysis including the *S*-tensor in order to account for correlations of libration and translation.¹¹ The results listed in Table 7 show that translational motion

Table 5. Observed [and unobserved with $F_u = \frac{1}{2}F_0(\text{min})$] and calculated structure factors on 10 times absolute scale.

Observed				Unobserved				Observed				Unobserved									
<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_u	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_u	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_u	<i>h</i>	<i>k</i>	<i>l</i>	F_o	F_u		
0	0	2	415	-428	0	5	3	44	40	1	1	7	47	50	1	4	-2	140	-161		
0	0	4	1283	-1322	0	5	4	150	136	1	1	9	248	272	1	4	-1	127	-125		
0	0	6	872	-916	0	5	5	39	32	1	1	11	282	-302	1	4	0	196	186		
0	0	8	140	139	0	5	6	111	-108	1	1	13	127	123	1	4	1	88	97		
0	0	10	154	163	0	5	8	9	0	1	1	15	148	149	1	4	2	236	-206		
0	0	12	190	203	0	5	10	27	24	1	1	17	35	-31	1	4	3	172	157		
0	0	14	39	29	0	5	11	84	84	1	1	19	141	-134	1	4	4	5	66	73	
0	0	16	67	-56	0	5	12	21	17	1	1	21	194	-172	1	4	6	139	149		
0	0	18	39	-40	0	5	13	51	45	1	1	23	68	50	1	4	7	70	-57		
0	0	20	18	39	0	5	14	27	23	1	1	25	17	-11	1	4	9	74	-74		
0	0	22	38	39	0	5	15	16	-19	1	1	27	42	-36	1	4	10	145	-118		
0	1	1	88	-58	0	5	16	20	20	1	1	29	51	41	1	4	11	31	-34		
0	1	2	776	792	0	5	17	18	-19	1	1	31	67	51	1	4	12	44	43		
0	1	3	602	-638	0	5	18	16	-14	1	1	33	74	48	1	4	13	71	-59		
0	1	4	26	24	0	5	19	15	-1	1	1	35	84	78	1	4	14	78	60		
0	1	5	131	135	0	6	1	40	40	1	2	1	92	92	1	4	15	96	98		
0	1	6	117	115	0	6	2	39	-36	1	2	2	-10	-56	-70	1	4	16	51	-38	
0	1	7	254	269	0	6	3	50	60	1	2	4	-18	20	-36	2	0	-22	30	27	
0	1	8	762	-780	0	6	4	17	20	1	2	6	-16	29	-32	2	0	-18	77	-80	
0	1	9	190	211	0	6	5	106	105	1	2	8	22	42	-155	2	0	-14	31	-34	
0	1	10	167	161	0	6	6	3	50	60	1	2	10	90	-107	2	0	-14	178	-182	
0	1	11	210	-210	0	6	7	36	39	1	2	12	-13	132	137	2	0	-12	210	220	
0	1	12	190	182	0	6	8	4	20	20	1	2	14	78	-64	2	0	-10	304	310	
0	1	13	50	50	0	6	9	1	20	20	1	2	16	113	-114	2	0	-8	186	-216	
0	1	14	138	129	0	6	10	10	20	20	1	2	18	158	-163	2	0	-6	386	370	
0	1	15	33	-27	0	6	11	35	-38	1	2	20	-8	117	109	2	0	-4	407	379	
0	1	16	88	-79	0	6	12	28	-28	1	2	22	-6	152	-179	2	0	-2	315	-313	
0	1	17	91	-87	0	6	13	51	48	1	2	24	116	109	1	2	0	-15	444	-458	
0	1	19	12	10	0	6	14	14	-7	1	2	26	-4	33	-29	2	0	2	309	222	
0	1	20	81	-68	0	6	15	20	-29	1	2	28	-3	443	435	2	0	4	212	-209	
0	1	21	1	-6	0	6	16	18	16	1	2	30	-2	36	-33	2	0	6	165	165	
0	1	22	34	-27	0	6	17	28	-20	1	2	32	-1	28	22	2	0	8	126	-163	
0	1	23	42	42	0	6	18	18	16	1	2	34	113	-116	2	0	10	40	66		
0	2	0	872	-893	0	7	2	28	-20	1	2	36	239	-275	2	0	12	61	63		
0	2	1	211	253	0	7	3	36	27	1	2	38	234	212	2	0	14	26	-28		
0	2	2	300	285	0	7	4	22	-20	1	2	40	116	109	2	0	16	47	50		
0	2	3	242	237	0	7	5	9	7	1	2	42	381	-357	2	0	20	37	-40		
0	2	4	233	209	0	7	6	12	-5	1	2	44	170	-177	2	0	23	25	-18		
0	2	5	360	342	0	7	7	48	40	1	2	46	204	192	2	0	25	127	120		
0	2	6	408	399	0	7	8	14	8	1	2	48	111	-83	2	0	27	80	63		
0	2	7	119	-131	0	7	9	32	24	1	2	50	84	-81	2	0	29	97	-95		
0	2	8	263	-253	0	7	10	9	-5	1	2	52	144	150	2	0	31	88	73		
0	2	9	115	-130	0	7	11	52	-50	1	2	54	73	83	2	0	33	129	-108		
0	2	10	184	80	0	7	12	23	-20	1	2	56	142	166	2	0	35	110	-98		
0	2	11	88	-91	0	8	0	59	45	1	2	58	99	99	2	0	37	14	42	36	
0	2	12	123	-149	0	8	1	13	12	1	2	60	109	-131	2	0	39	13	78	60	
0	2	13	116	-121	0	8	2	19	12	1	2	62	156	-195	2	0	41	27	22		
0	2	14	89	92	0	8	3	36	27	1	2	64	146	-32	2	0	43	181	154		
0	2	15	210	222	0	8	4	44	-41	1	2	66	177	64	2	0	45	161	151		
0	2	16	73	80	0	8	5	36	-37	1	2	68	12	4	2	0	-9	169	-176		
0	2	17	19	-24	0	8	6	44	12	1	2	70	21	9	22	1	-8	18	-29		
0	2	18	40	49	0	8	7	25	-13	1	2	72	22	11	13	2	-7	86	-91		
0	2	19	8	-9	0	8	8	10	25	1	3	1	-7	22	-23	2	0	-6	106	-128	
0	2	20	36	-42	1	0	-22	11	-19	1	3	3	-19	13	-13	2	1	-5	102	-113	
0	2	21	41	-50	1	0	-18	74	70	1	3	5	-16	104	88	2	1	-4	151	-171	
0	2	22	90	-31	1	0	-16	26	-26	1	3	7	-15	64	-66	2	1	-4	152	174	
0	3	1	275	-298	1	0	-14	88	68	1	3	9	-14	224	-220	2	1	-2	178	-165	
0	3	2	90	85	1	0	-12	150	129	1	3	11	-13	128	112	2	1	-1	102	109	
0	3	3	56	-65	1	0	-10	761	-804	1	3	13	-11	170	168	2	1	0	41	-29	
0	3	4	516	-518	1	0	-8	185	-191	1	3	15	-9	299	31	2	1	0	40	-25	
0	3	5	277	269	1	0	-6	1013	1069	1	3	17	-8	210	211	2	1	1	49	45	
0	3	6	49	42	1	0	-4	7048	1087	1	3	19	-7	318	-307	2	1	4	103	-111	
0	3	7	97	99	1	0	0	964	-1000	1	3	21	-6	432	-417	2	1	5	110	-127	
0	3	8	200	204	1	0	0	6	239	269	1	3	23	-5	53	-52	2	1	6	75	72
0	3	9	144	-151	1	0	1	4	927	-950	1	3	25	-4	65	-64	2	1	7	38	-34
0	3	10	76	64	1	0	2	3	457	470	1	3	27	-3	80	-77	2	1	8	84	-87
0	3	11	156	156	1	0	3	8	258	293	1	3	29	-2	95	-92	2	1	9	92	87
0	3	12	61	54	1	0	4	12	244	-245	1	3	31	-1	108	-106	2	1	10	109	103
0	3	13	145	-134	1	0	5	16	88	47	1	3	33	0	255	-268	2	1	11	171	-166
0	3	14	50	-40	1	0	6	20	17	15	1	3	35	2	275	275	2	1	12	154	155
0	3	15	52	-49	1	0	7	22	50	-44	1	3	37	1	76	64	2	1	13	151	130
0	3	16	54	54	1	0	8	22	50	-44	1	3	39	0	158	-172	2	1	14	109	95
0	3	17	72	64	1	0	9	22	50	-44	1	3	41	0	289	-279	2	1	15	136	-114
0	3	18	33	-28	1	0	10	22	50	-44	1	3	43	0	411	123	2	1	16	142	-121
0	3	19	14	8	1	0	11	22	50	-44	1	3	45	0	117	114	2	1	17	37	-32
0	3	20	20	14	1	0	12	22	50	-44	1	3	47	0	289	-279	2	1	20	27	22
0	3	21	6	5	1	0	13	22	50	-44	1	3	49	0	411	123	2	1	21	22	17
0	4	0	236	-212	1	1	1	-19	66	-67	1	3	51	0	188	-194	2	2	-10	14	27
0	4	1	94	118	1	1	2	-18	40	32	1	3	53	0	33	30	2	2	-16	32	-41
0	4	2	97	108	1	1	3	-16	77	-58	1	3	55	0	147	137	2	2	-18	32	-41
0	4	3	150	-184	1	1	4	-13	100	94	1	3	57	0	177	84	2	2	-20	59	27
0	4	4	59	-68	1	1	5	-12	48	48	1	3	59	0	135	-128	2	2	-22	81	-82
0	4	5	98	-97	1	1	6	-11	55	47	1	3	61	0	127						

Table 5. Continued.

Table with multiple columns of numerical data, organized into groups by header letters h, k, l. The data consists of three main columns of values for each group, representing different parameters or conditions.

Table 5. Continued.

k	l	F _o	F _c	h	k	l	F _o	F _c	h	k	l	F _o	F _c	h	k	l	F _o	F _c		
5	1	5	91	-75	6	0	-10	431	380	6	3	5	110	-118	7	2	5	173	-179	
5	1	6	193	-193	6	0	-8	159	152	6	3	6	101	-110	7	2	6	86	-88	
5	1	7	244	252	6	0	-6	20	39	6	3	7	99	-87	7	2	7	80	-40	
5	1	8	187	-192	6	0	-4	223	-200	6	3	8	69	-64	7	2	8	38	-41	
5	1	9	65	-59	6	0	-2	97	-101	6	3	9	69	-63	7	2	9	42	-50	
5	1	10	183	90	6	0	0	138	-154	6	3	10	47	34	7	2	10	43	-51	
5	1	11	20	19	6	0	2	373	405	6	3	11	24	28	7	2	11	34	-4	
5	1	12	114	103	6	0	4	66	-76	6	3	12	26	29	7	2	12	27	-38	
5	1	13	11	-2	6	0	6	87	87	6	3	13	31	30	7	2	13	38	72	
5	1	14	42	-34	6	0	8	101	-95	6	3	14	27	72	7	2	14	22	17	
5	1	15	38	-74	6	0	10	104	-115	6	4	-15	52	-67	7	3	-18	22	70	
5	1	16	63	-49	6	0	12	29	-23	6	4	-14	31	-60	7	3	-17	37	-32	
5	1	17	14	17	6	0	14	110	99	6	4	-13	100	-117	7	3	-15	40	-43	
5	1	18	26	-24	6	1	-23	22	-17	6	4	-10	108	-120	7	3	-14	14	-3	
5	2	-10	137	-61	6	1	-22	55	-48	6	4	-9	148	-173	7	3	-13	143	-123	
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Table 5. Continued.

h	k	l	F _o	F _c	h	k	l	F _o	F _c	h	k	l	F _o	F _c	h	k	l	F _o	F _c
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8	2	-8	136	- 140	9	1	2	128	- 121	10	2	-10	116	- 99	12	0	0	68	56
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8	2	-5	374	345	9	1	5	108	- 94	10	2	-7	24	- 34	12	1	-14	23	16
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8	2	-1	57	- 48	9	1	9	49	- 41	10	2	-2	28	- 19	12	1	-10	9	- 4
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8	4	25	11	- 25	10	1	-14	60	- 58	11	1	16	14	- 10					
8	4																		

Table 5. Continued.

h	k	l	F _h	F _c	h	k	l	F _h	F _c	h	k	l	F _h	F _c	h	k	l	F _h	F _c
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2	2	-17	8	19	5	0	2	6	8	7	4	13	17	-11	10	2	-19	4	-13
2	2	14	8	29	5	0	16	7	-8	7	4	-7	17	20	10	2	-17	6	-7
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2	2	21	4	11	5	2	-21	7	6	8	0	-22	5	-3	10	2	5	5	0
2	3	-10	8	-9	5	2	-19	7	-7	8	0	-16	7	-12	10	2	8	5	-8
2	3	12	8	16	5	2	-11	7	-2	8	0	6	8	8	10	2	9	4	9
2	3	15	8	11	5	2	-8	6	-12	8	0	10	7	3	10	3	-19	3	-13
2	3	17	7	2	5	2	16	6	12	8	1	-21	4	11	10	3	-17	5	5
2	4	-17	16	-5	5	2	14	4	-23	8	1	-20	5	8	10	3	-16	6	7
2	4	-19	17	-11	5	3	-22	5	-9	8	1	-9	6	6	10	3	-15	6	4
2	4	-2	17	-11	5	3	10	7	-11	8	1	-7	6	-2	10	3	-7	7	-14
2	4	-8	14	-11	5	3	6	7	-7	8	1	5	6	-2	10	3	-2	7	-8
2	4	-6	12	-31	5	3	13	7	3	8	1	12	5	2	10	3	1	7	-8
2	4	-4	10	-29	5	3	5	5	-11	8	1	14	3	6	10	3	6	5	-8
2	4	16	14	-27	5	4	-17	16	-11	8	0	10	7	3	10	3	7	4	-23
2	4	10	17	-12	5	4	-12	17	-11	8	2	-19	6	-12	10	4	-6	15	-9
2	4	11	17	-10	5	4	-9	16	12	8	2	-11	8	-8	11	0	-12	7	-8
3	0	-22	6	-12	5	4	-1	15	-28	8	2	6	8	-5	11	0	0	7	-2
3	0	-20	8	-13	5	4	-1	15	-28	8	2	6	8	-5	11	0	4	7	-13
3	0	16	8	0	5	4	8	17	24	8	3	-14	8	6	11	1	-12	5	-10
3	1	-23	8	6	5	4	13	15	-20	8	3	4	8	15	11	1	-11	6	-5
3	1	-22	5	-5	6	0	-22	6	-16	8	3	12	4	13	11	1	-9	6	9
3	1	-17	6	-5	6	0	-18	8	-10	8	3	-15	15	7	11	1	-6	6	10
3	1	-14	6	-1	6	0	16	6	3	8	4	-13	16	9	11	1	-1	6	15
3	1	14	6	-4	6	0	14	5	-3	8	4	-11	17	10	11	1	4	4	0
3	1	21	6	-4	6	0	-21	7	6	8	4	-10	17	20	11	2	-18	3	-6
3	2	18	6	-20	6	2	-20	7	-11	8	4	-10	17	34	11	3	-17	7	-9
3	2	20	4	5	6	2	-18	7	-6	8	4	-6	17	-11	11	2	-15	5	3
3	3	-20	7	10	6	2	-14	8	-12	9	0	8	7	0	11	2	-14	6	35
3	3	-17	8	9	6	2	-12	7	9	9	1	-19	5	9	11	2	-12	6	12
3	3	18	6	-7	6	2	-11	7	3	9	1	11	4	-3	11	2	-10	7	-11
3	4	-17	10	-18	6	2	12	7	-15	9	2	-18	6	-11	11	2	-8	7	-2
3	4	-16	17	-18	6	2	16	5	-31	9	2	-18	6	-11	11	2	-6	7	-10
3	4	-9	15	-16	6	3	-22	4	-12	9	2	-3	8	1	11	2	-3	7	-9
3	4	-5	12	-24	6	3	-22	8	31	11	2	1	8	6	11	2	2	7	0
3	4	-4	11	-2	6	3	15	5	-1	9	2	5	7	16	11	2	4	5	-6
3	4	-3	11	-1	6	4	-17	15	-11	9	2	8	6	14	11	2	5	4	-16
4	0	-22	6	-4	6	4	-16	16	-16	9	2	10	5	4	11	2	-6	4	-19
4	0	-20	6	-6	6	4	-12	17	-16	9	2	11	4	-20	11	3	-7	5	-10
4	0	18	6	-6	6	4	-11	17	-12	9	3	-20	6	-3	11	3	-14	5	-13
4	1	-23	4	2	6	4	8	17	-14	9	3	-19	5	14	11	3	-12	6	6
4	1	-14	6	0	6	4	10	16	-4	9	3	-16	6	3	11	3	-6	7	7
4	1	-5	4	6	6	4	11	15	-34	9	3	-14	6	2	12	1	-16	7	-9
4	1	8	6	7	7	0	-16	8	-6	9	3	-4	8	-7	12	1	-3	5	-9
4	2	-21	6	-1	7	0	4	8	14	9	3	2	8	-9	12	1	-2	5	6
4	2	-14	8	-9	7	0	16	4	11	9	3	6	7	23	12	1	-1	4	4
4	2	-9	6	-8	7	2	-8	7	7	9	4	-11	4	-1	12	1	-3	3	-9
4	2	8	6	8	7	2	-21	5	8	9	4	-12	15	-20	12	2	-13	5	3
4	2	14	8	7	7	2	-17	8	-12	9	4	-9	17	-27	12	2	-12	5	-0
4	2	16	7	9	7	2	-10	8	-12	9	4	-7	17	40	12	2	-5	6	11
4	2	18	4	-5	7	2	-4	7	-2	9	4	-6	17	1	12	2	-10	5	-2
4	2	19	4	-15	7	2	-4	7	-2	9	4	3	15	-25	12	3	-14	3	10
4	3	-12	7	-6	7	2	0	7	9	10	0	-20	4	5	12	3	-13	3	19
4	3	-6	6	-12	7	2	13	6	-27	10	0	-8	8	16	12	3	-8	5	5
4	3	8	7	-20	7	2	15	4	9	10	0	-6	8	-3	12	3	-6	5	2
4	3	11	8	11	7	1	-22	3	4	10	0	6	6	-4	12	0	6	3	-2
4	3	12	8	18	7	3	-21	5	9	10	0	10	4	-11	13	0	-6	5	5
4	4	-17	16	-5	7	3	-20	6	-5	10	1	-15	6	-9	13	0	-4	4	-2
4	4	-16	17	11	7	3	-19	6	-16	10	1	-5	6	-9	13	1	-9	3	-2
4	4	-13	17	12	7	3	-10	8	-25	10	1	-2	6	-21	13	2	-7	3	16
4	4	-12	17	38	7	3	-9	8	10	10	1	-1	6	-0					
4	4	-11	17	15	7	3	7	8	8	10	1	0	6	-0					

Table 6. The principal axes of the thermal vibration ellipsoids given by the components of a unit vector in fractional coordinates e_x , e_y , e_z ; the corresponding r.m.s. amplitudes, and the B -values.

Atom	e_x	e_y	e_z	$(\bar{u}^2)^{1/2}(\text{Å})$	$B(\text{Å}^2)$
O ₁	.005	.045	-.049	.283	6.31
	.048	-.088	.005	.232	4.26
	.063	.104	.025	.205	3.32
O ₂	.043	.023	.054	.240	4.53
	.010	-.142	.008	.221	3.85
	.090	.006	-.008	.207	3.38
O ₃	.029	.070	.048	.256	5.16
	.016	-.125	.027	.210	3.49
	-.095	.000	.002	.178	2.49

Table 6. Continued.

O ₄	.017	-.035	.053	.252	5.02
	.027	.136	.015	.207	3.37
	.095	-.032	.004	.178	2.51
O ₅	.013	.060	-.045	.226	4.04
	.020	.126	.026	.221	3.84
	.098	-.033	.018	.196	3.04
O ₆	.037	.112	.032	.298	6.99
	.080	-.078	.027	.219	3.79
	.048	.044	-.035	.206	3.36
C ₁	.044	.073	.046	.234	4.28
	.035	.094	-.029	.219	3.77
	.083	-.079	.008	.183	2.64
C ₂	.053	.059	.050	.294	6.84
	.011	.126	-.023	.218	3.77
	.093	-.037	.002	.183	2.65
C ₃	.027	.065	.049	.292	6.74
	.003	.126	-.025	.243	4.65
	.097	-.022	.004	.189	2.82
C ₄	0.46	.080	.044	.274	5.92
	.031	.090	-.032	.256	5.19
	.083	-.078	.008	.192	2.90
C ₅	.027	.015	.055	.288	6.54
	.075	.087	-.002	.241	4.57
	.061	-.113	.006	.194	2.98
C ₆	.034	-.025	.054	.267	5.63
	.082	.077	.003	.209	3.45
	-.046	.119	.009	.191	2.88
C ₇	.002	-.028	.052	.236	4.41
	.035	.132	.015	.198	3.10
	.094	-.049	.011	.178	2.51
C ₈	.013	-.047	.051	.264	5.49
	.003	.135	.018	.225	4.01
	.100	.002	.010	.180	2.57
C ₉	.020	.094	.041	.287	6.52
	.001	-.106	.035	.254	5.11
	.098	-.018	.008	.186	2.73
C ₁₀	.035	.084	.044	.290	6.66
	.028	.098	-.031	.250	4.94
	.090	-.063	.011	.177	2.47
C ₁₁	.017	-.008	.054	.261	5.37
	.057	.118	.007	.235	4.35
	.081	-.081	.005	.191	2.88
C ₁₂	.020	.007	.055	.245	4.73
	.007	.143	-.002	.208	3.43
	.098	-.012	.006	.193	2.95

Table 7. Results of the rigid-body analysis.

a) In the orthogonal system, J_1, J_2, J_3 , defined by $J_1||a, J_2||b, J_3||c^*$. (E.s.d. of components of L are given in parentheses in units of last place shown).

$$L = \left\{ \begin{array}{ccc} 163 (16) & -95 (11) & -44 (7) \\ & 101 (9) & 43 (7) \\ & & 62 (5) \end{array} \right\} \times 10^{-1} (^\circ)^2$$

	r.m.s. amplitude	Direction cosines ($\times 10^3$)		
L	5.0°	J_1 761	J_2 -572	J_3 306
	2.2°	520	257	-814
	1.6°	387	779	494
T^a	0.23 \AA	-069	-038	997
	0.20 \AA	199	-980	-023
	0.18 \AA	978	197	075

b) In the orthogonal system, L_1, L_2, L_3 , defined by the principal axes of the libration tensor L .

Displacement of libration axes from intersecting (\AA)

Parallel to L_1	-0.19
» » L_2	-0.67
» » L_3	0.11

Effective screw translations (\AA)

Parallel to L_1	0.004
» » L_2	-0.014
» » L_3	0.008

$$(\overline{\Delta U_{ij}^2})^{\frac{1}{2}} = 0.0035 \text{ \AA}^2$$

^a T is the reduced translation tensor.

Table 8. Corrections in fractional atomic coordinates due to libration compared with estimated standard deviations in coordinates. (σ 's and Δ 's are multiplied by 10^6).

Atom	Δx	$\sigma(x)$	Δy	$\sigma(y)$	Δz	(z)
O ₁	-58	28	-83	53	-41	17
O ₂	-25	26	-3	47	-26	15
O ₃	-15	23	-16	47	22	15
O ₄	25	22	40	46	21	14
O ₅	6	23	-31	45	-28	13
O ₆	29	27	29	53	-44	16
C ₁	-22	34	8	69	9	20
C ₂	-62	38	-43	82	16	28
C ₃	-72	37	-21	85	6	27
C ₄	-40	42	61	86	21	27
C ₅	-2	44	111	87	11	29
C ₆	10	39	91	74	20	25
C ₇	25	31	0	67	5	20
C ₈	68	33	60	78	10	24
C ₉	72	38	24	86	-3	27
C ₁₀	49	38	-40	86	12	26
C ₁₁	5	37	-101	81	5	25
C ₁₂	0	35	-68	71	17	22

is reasonably isotropic, while libration is more anisotropic, with the primary motion corresponding to an r.m.s. amplitude of 5.0° . The r.m.s. discrepancy between observed and calculated U_{ij} serves as a measure of the applicability of the analysis and thus of the rigidity of the molecule. The value of Table 7, 0.0035 \AA^2 , supports the assumption of regarding the molecule as an oscillating rigid body, and the fractional coordinates were accordingly corrected for librational effects.¹² Table 8 shows that all corrections are less than three times the corresponding standard deviations of the coordinates.

DISCUSSION

Interatomic distances and angles (corrected for librational effects) may be found in Figs. 1, 2, and 3. The standard deviations, given in parentheses (in units of last place shown), have been calculated from the correlation matrix corresponding to the last full-matrix least squares refinement cycle without taking into account the e.s.d.'s of the cell parameters. The largest correction

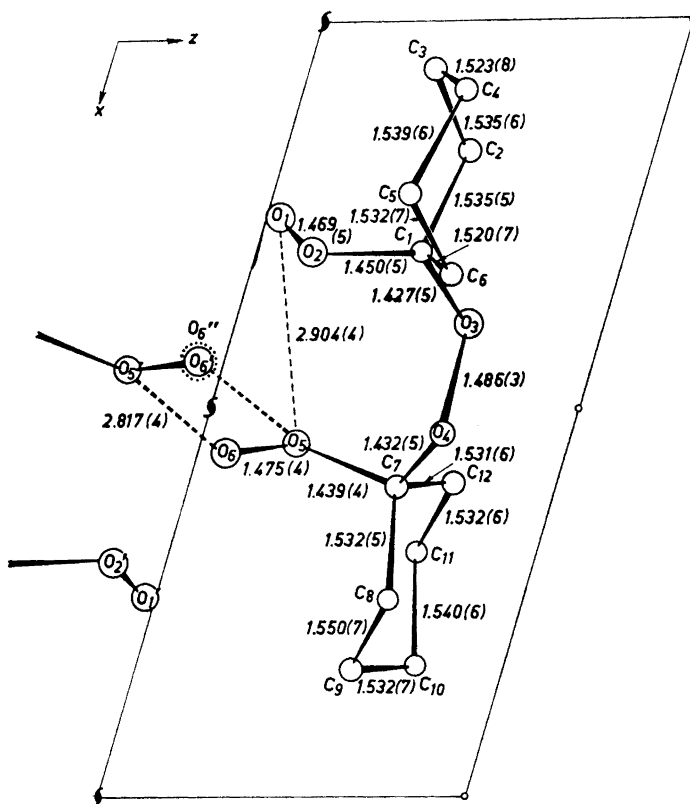


Fig. 1. Schematic drawing of the molecule showing interatomic distances.

for libration in distances and angles are $1.6 \times \sigma_d$ and $0.6 \times \sigma_a$, where σ_d and σ_a are the corresponding standard deviations. The C—H distances arrived at range from 0.91 Å to 1.14 Å with estimated standard deviations of about 0.06 Å.

From Fig. 1 it may be seen that the three O—O bond distances as well as the four C—O bond lengths, respectively, are equal within probable limits of error. The corresponding mean values, 1.47₇ Å and 1.43₇ Å, agree closely with earlier findings.¹³⁻¹⁷ The C—C bond lengths are also normal, with mean value 1.53₃ Å.

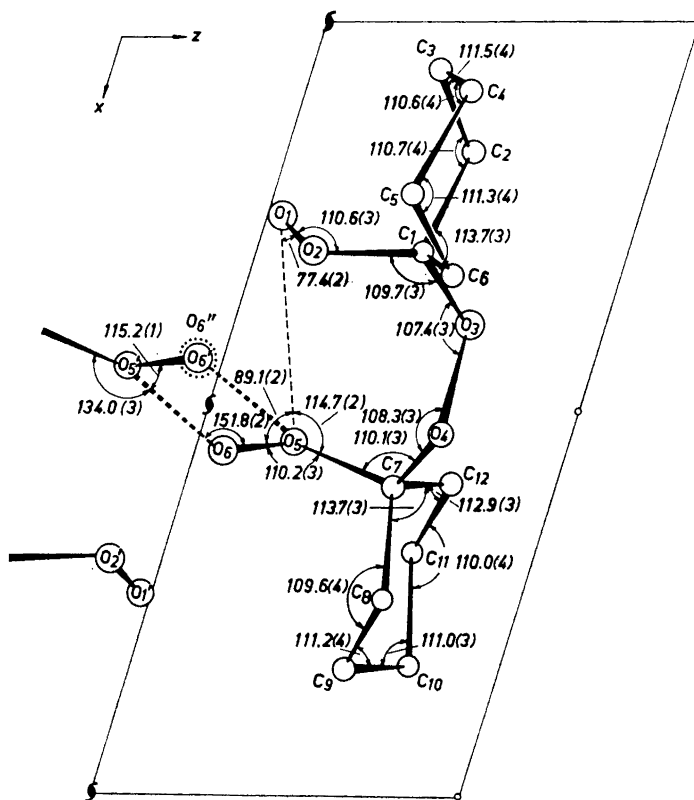


Fig. 2. Schematical drawing of the molecule showing bond angles.

The angles (Fig. 2) $C_1-O_2-O_1$ (110.6°) and $C_7-O_5-O_6$ (110.2°) are equal and significantly larger than $C_1-O_3-O_4$ (107.4°) and $C_7-O_4-O_3$ (108.3°). The larger C—O—O angles correspond to the values obtained for dibenzoyl peroxide⁴ (109° and 111°), while the smaller agree with those found in trimeric and dimeric peroxides.¹³⁻¹⁷

The arrangements around the carbon atoms C_1 and C_7 are asymmetric as shown in Fig. 3 (a) and (b). The angles $C_2-C_1-O_3$ (103.6°), $C_6-C_1-O_2$

(104.0°), $C_8-C_7-O_4$ (103.2°), and $C_{12}-C_7-O_5$ (104.3°) are considerably smaller, (about 9° in average), than $C_2-C_1-O_2$ (111.8°), $C_7-C_1-O_3$ (114.2°), $C_8-C_7-O_5$ (112.9°), and $C_{12}-C_7-O_4$ (112.9°). The kind of asymmetry resembles what was observed for trimeric acetone peroxide,¹⁷ and the distortions may roughly be described as "rotations" (*ca.* 6°) about axes through C_1

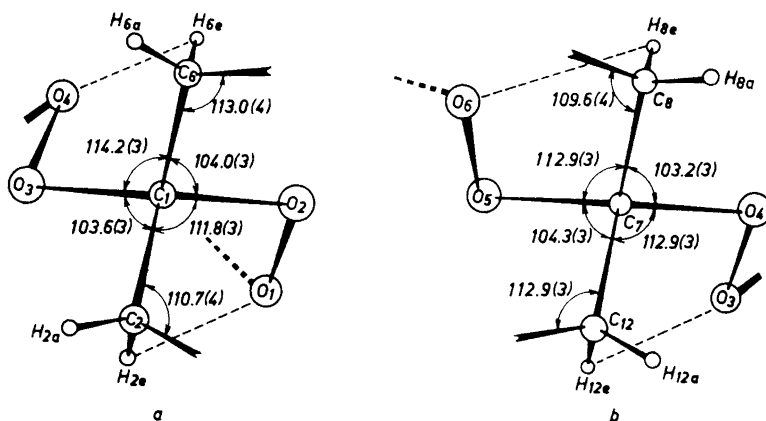


Fig. 3. Schematical drawings showing bond angles and the asymmetric environments of the carbon atoms C_1 and C_7 .

and C_7 normal to the paper plane (Fig. 3). In dimeric peroxides the corresponding axes of "rotation" are normal to the C—C—C planes.

The cyclohexane rings are also slightly distorted; the angles $C_1-C_2-C_3$ (110.7°) and $C_7-C_8-C_9$ (109.6°) being possibly significantly smaller than $C_1-C_6-C_5$ (113.0°) and $C_7-C_{12}-C_{11}$ (112.9°). As pointed out elsewhere,^{13,17} it seems reasonable to relate the effects to *intra*-molecular oxygen-hydrogen repulsions (see Fig. 3). The angles $C_2-C_1-C_6$ and $C_8-C_7-C_{12}$ are both equal to 113.7° while the rest of the C—C—C angles have normal cyclohexane values (110.8° in average with standard deviation 0.3°).

The conformation of the molecule with respect to the hydroperoxy groups is *cis*, and the dihedral angle $C_1-O_3-O_4-C_7$ is found to be 126.3° (3), which is 35° larger than the dibenzoyl peroxide value.⁴

As mentioned above, the hydroperoxy hydrogen atoms could not be localized in the difference Fourier map. The solid state infrared spectrum (nujol) recorded in the region 3200 cm^{-1} to 3600 cm^{-1} , containing two bands, at 3385 cm^{-1} and 3420 cm^{-1} , predicts two different kinds of hydrogen bonds, since ν_{OH} for hydroperoxides should be in the region 3530 cm^{-1} to 3560 cm^{-1} .¹⁸ In Fig. 1 the dotted lines from O_6 to O_5' , and from O_5 to O_6'' in a neighbouring cell, indicate *inter*-molecular hydrogen bonds of length 2.81 \AA , forming helical chains in direction of the *b*-axis. Since the shortest *inter*-molecular O_1-O distance, O_1-O_6'' , is about 3.42 \AA , the hydrogen bond associated with the O_2O_1H -group has to be *intra*-molecular. The distance O_1-O_5 and O_1-O_6 of lengths 2.90_4 \AA and 3.23_8 \AA , respectively, together with the corre-

sponding angles, $O_2-O_1-O_5$ (77.4°) and $O_2-O_1-O_6$ (62.1°), point out O_1-O_5 as a weak *intra*-molecular hydrogen bond. The extreme values of the $O-O-O$ angles, (77.4° and 151.8°), show that the hydrogen bonds in this compound are far from linear. The angles of bond direction at the oxygen atom O_5 , which are accepting *both* the hydrogens, are closer to normal values: $O_1-O_5-O_6$ (89.1°), $O_1-O_5-C_7$ (114.7°), $O_6''-O_5-O_6$ (115.2°), and $O_6''-O_5-C_7$ (134.0°).

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